Chemistry 115 Name

Dr. Cary Willard

Quiz 5E October 8, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of sodium sulfate, Na2SO4.
2. (4 points) Calculate the mass in grams of 4.52 mol of sodium sulfate.
3. (4 points) Calculate the mass in grams of 1.85 x 1024 atoms of silver.
4. (4 points) Calculate the number of mols of oxygen in 6.27 mols of sodium sulfate.
5. (4 points) Determine the empirical formula of a compound that is 43.41% W and 56.59% Br

Chemistry 115 Name

Dr. Cary Willard

Quiz 5F October 8, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of sodium sulfate, Na2SO4.
2. (4 points) Calculate the mass in grams of 6.32 mol of sodium sulfate.
3. (4 points) Calculate the mass in grams of 3.65 x 1024 atoms of silver.
4. (4 points) Calculate the number of mols of oxygen in 3.22 mols of sodium sulfate.
5. (4 points) Determine the empirical formula of a compound that is 31.52% W and 68.48% Br

Chemistry 115 Name

Dr. Cary Willard

Quiz 5A October 6, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of lithium phosphite, Li3PO3.
2. (4 points) Calculate the number of moles of lithium phosphite in 85.3 g of lithium phosphite.
3. (4 points) Calculate the number of atoms of titanium in a 4.00 g sample of titanium.
4. (4 points) Calculate the number of mols of oxygen in 4.97 mols of lithium phosphite.
5. (4 points) Determine the empirical formula of a compound that is 21.94% S and 78.03% F

Chemistry 115 Name

Dr. Cary Willard

Quiz 5B October 6, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of lithium phosphite, Li3PO3.
2. (4 points) Calculate the number of moles of lithium phosphite in 36.5 g of lithium phosphite.
3. (4 points) Calculate the number of atoms of titanium in a 3.00 g sample of titanium.
4. (4 points) Calculate the number of mols of oxygen in 6.72 mols of lithium phosphite.
5. (4 points) Determine the empirical formula of a compound that is 24.59% P and 74.41% F

Chemistry 115 Name

Dr. Cary Willard

Quiz 5C October 7, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of potassium carbonate, K2CO3.
2. (4 points) Calculate the number of moles of potassium carbonate in 6.86 g of potassium carbonate.
3. (4 points) Calculate the number of atoms of chromium in a 6.44 g sample of chromium.
4. (4 points) Calculate the number of mols of oxygen in 2.53 mols of potassium carbonate.
5. (4 points) Determine the empirical formula of a compound that is 62.61% Sn and 37.39% Cl

Chemistry 115 Name

Dr. Cary Willard

Quiz 5D October 7, 2009

NA = 6.022 x 1023/mol

1. (4 points) Determine the molar mass of potassium carbonate, K2CO3.
2. (4 points) Calculate the number of moles of potassium carbonate in 2.89 g of potassium carbonate.
3. (4 points) Calculate the number of atoms of chromium in a 4.86 g sample of chromium.
4. (4 points) Calculate the number of mols of oxygen in 8.55 mols of potassium carbonate.
5. (4 points) Determine the empirical formula of a compound that is 45.56% Sn and 54.43% Cl